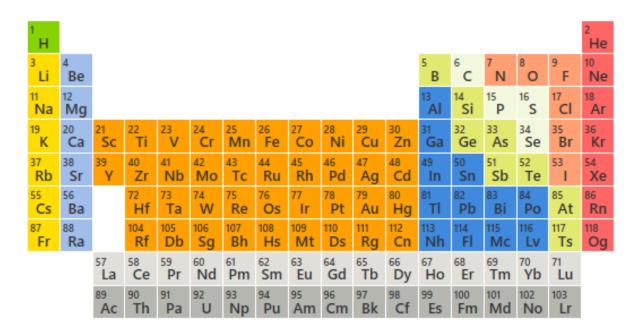
Wordsworth High School Grade 9 NS Chemistry

RECAP Periodic Table



Group 1: Alkali Metals

- Only 1 electron in outer orbit

Group 2: Alkali-earth Metals

- Only 2 electrons in outer orbit

Group 17: Halogens

- Only 1 electron space in outer orbit

Group 18: Noble Gases (stable)

- all electrons in outer orbit is full.

Chemical Reactions.

Chemical Reactions are represented by chemical equations

• Chemical reactions can be represented with models.



 $C+O_2 \rightarrow CO_2$

 $2H_2 + O_2 \rightarrow 2H_2$

Chemical reactions can also be represented by symbols if it occurs in a balanced equation, eg.

 $C+O_2 \rightarrow CO_2$

$2H_2 + O_2 \rightarrow 2H_2O$

The small two (undercase) indicates to us the amount of atoms in the compound.

The big two in front of hydrogen, shows the relationship of the chemical reaction. Eg. 2 molecules Hydrogen reacts whith 1 molecule Oxygen to form water. Therefore the equation will be 2:1 (H:O).

Matter can not be made or destroyed.

Atoms can only rearrange.

All the atoms infront of the arrow are called reactants. The atoms behind the arrow are called products.